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Therefore The Ratio Is 1:1, So The Empirical Formula = HF B. $(4.267g\ UF6\ X\ 1\ Mol\ UF6\ X\ 6\ Mol\ F)/(352.03g\ X\ 1mol) = 0.07273\ Mol\ F\ Original\ Compound\ Gas: (0.0485\ Mol\ F)/(0.07273\ Mol\ F) = 66.7\%\ In\ Gas\ Or\ 922/1382\ Solid: 33.3\%\ In\ Solid\ Or\ 460/1382\ C. (4.267g\ UF6\ X\ 1\ Mol\ UF6\ X\ 1\ Mar\ 5th,\ 2024$

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