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Key Channels Identified Via Detailed Path Analysis And With Rates Determined  
Using . Ab Initio . ... -1 Enthalpy Of Formation 2 Apr 18th, 2024CHEMICAL KINETICS -  
VSSUT2and Cl 2to FormHCl. 7 Chemistry-I. Figure 1.1: Plot Of Concentration Against  
T For A Reaction  $H_2 + Cl_2 \rightarrow 2HCl$  The Rate Of This Reaction Does Not Depend On  
Concentration. 1.6 First Order Kinetics A Reaction Of The first Order Is Represented  
As  $X \rightarrow Y$  Where X Is The Reactant And Y The Product. The Rate Of The Reaction Will  
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Transfer Translational KE To Vibrational Energy To Weaken/break Reactant Bonds  
Molecules At A Given Temperature Possess A KE Distribution: Number Of Molecules  
KE  $E \geq E_A$ , Activation Energy - Energy Barrier That Molecules Have To Surmount In  
Order To React. Energy Is Needed To Break Reactant Bonds (endothermic Process).  
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Exactly Is Reacting? Usually Ions Of Opposite Charge React Very Rapidly. Also, The  
More Bonds Between Reacting Atoms In A Molecule, The Slower The Reaction Rate.  
Why? More Energy Is Required To Separate The Molecule Into Its "bits". Substances  
With Strong Bonds (larger Bond Energies) Will React Much More Slowly. Mar 3th,  
2024.

Chapter 14 Chemical Kinetics3 Chemical Kinetics Factors That Affect Reaction Rates  
• Physical State Of The Reactants In Order To React, Molecules Must Come In  
Contact With Each Other. If The Reaction Is Happening Between A Solid And A Liquid  
It Will React Only On The Surface. The More Homogeneous The Mixture Of  
Reactants, The Faster The Molecules Can React. Apr 9th, 2024

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