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WARM-UP (1) Convert 78°C To Kelvin 351 K (2) $3.45\text{ atm} = \underline{\hspace{1cm}}\text{ mmHg}$ 2622 mmHg (3) The Coldest Temperature Possible Is (0 K) = Absolute Zero (4) What Is The Final Volume Of A 400.0 mL Gas Sample That Is Subjected To A Temperature Change From 22.0°C To 30.0°C And A Pressure Change From 760.0 mm Hg To 360.0 mm Hg ? REMEMBER: Temp Must Be In Kelvin!! Mar 19th, 2024

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Container? 1. The He Gas Because It Is Less Massive And Moving With A Higher Average Velocity Correct 2. They Are The
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