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4) -A Solution Is Created By Measuring 3.60×10^3 Moles Of NaOH And 5.95×10^{-4} Moles Of HCl Into A Container And Then Water Is Added Until The Final Volume Is 1.00 L. What Is The PH Of This Solution? Since There Is Both Acid And Base We Will Assume A 1 Mole Acid:1 Mole Base Ratio Of Neutralization. Apr 9th, 2024

PH And POH - Ms. Mogck's Classroom - Ms. Mogck's Classroom

KEY Chemistry: PH And POH Calculations Part 1: Fill In The Missing Information In The Table Below. PH [H 3 O 1+] POH [OH1-] ACID Or BASE? 3.78 1.66×10^{-4} M 10.22 6.03×10^{-11} M Acid 3.41 3.89×10^{-4} M 10.59 2.57×10^{-11} M Acid 8.81 1.55×10^{-9} M 5.19 6.46×10^{-6} M Base 8.69 2.04×10^{-9} M 5.31 4.88×10^{-6} M Base 8.46 3.47×10^{-9} M 5.54 2.88×10^{-6} M Base Feb 21th, 2024

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