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11 Stoichiometry Review Worksheet Name: ___KEY___ 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$

Actual yield = 1.45g Theoretical yield = 5.95g $\text{KClO}_4 \times 1 \text{ mol } 122.55 \text{ g} \times 2 \text{ O}_2 \times 1 \text{ KClO}_4 \times$

$32.00 \text{ g } 1 \text{ mol} = 3.11 \text{ g O}_2$ %Yield = $\frac{1.45 \text{ g } 3.11 \text{ g}}{3.11 \text{ g}} \times 100\% = 46.6\%$ 34. $4\text{FeCO}_3 + \text{O}_2$

$\rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$ Percent purity = 62.8%? g pure FeCO_3 6th, 2024.

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Of LiNO_3) 4) Using The Following Equation: $\text{Fe}_2\text{O}_3 + 3 \text{H}_2 \rightarrow 2 \text{Fe} + 3 \text{H}_2\text{O}$.

Calculate How Many Grams Of Iron Can Be Made From 16.5 Grams Of Fe_2O_3 By The

Following Equation. 23th, 2024Basic Stoichiometry Phet Post Lab Answer KeyBasic

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Bing Phet Simulation The Skater Had More Potential Energy At The Bottom And Had

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