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Of ...  $\Delta G^\circ = -RT \ln K$  Standard Emf Of An

Electrochemical Cell:  $E^\circ_{\text{Cell}} = E^\circ_{\text{Ox}} - E^\circ_{\text{Red}} = E^\circ_{\text{Cathode}} - E^\circ_{\text{Anode}}$  Standard Free Energy Change:  $\Delta G^\circ$

=  $-nFE^\circ_{\text{Cell}}$  Where F Is The Faraday Constant  
Relationship Of The Standard Emf Of The Cell To The

Equilibrium Constant:  $E^\circ_{\text{Cell}} = \frac{RT}{nF} \ln K$  The Nernst  
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