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Nitrogen ( N<sub>2</sub>) 100 150 200 250 300 350 450 550 700 1,000 1,100 1,200 1,300

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Follows That The Sum Of All The Mole Fractions Of The Mixture Must Also Equal 1.  
 $X_1 + X_2 + \dots = 1$  The Mass Of Component I In A Mixture Is The Product Of The  
Number Of Moles Of I And Its Molecular Weight,  $M_i$ . The Mass Of The Mixture Is  
Therefore The Sum,  $M = N_1M_1 + N_2M_2 + \dots$  Jan 11th, 2024 Combustion Turns Solution  
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