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33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$  Actual yield = 1.45g  
Theoretical yield =  $5.95\text{g KClO}_4 \times \frac{1\text{mol}}{122.55\text{g}} \times 2 \times \frac{1\text{mol O}_2}{1\text{mol KClO}_4} \times 32.00\text{g} = 3.11\text{g O}_2$   
% Yield =  $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$

34.  $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$   
Percent purity = 62.8%  
2024Chemistry 12, Stoichiometry Review Problems, R.  
Janssen ... In A Camping Stove? 2. How Many Moles Of  
 $\text{O}_2$  Molecules Should Be Supplied To Burn 1 Mol Of  $\text{CH}_4$  ...  
Was At One Time Considered For Use As A Rocket  
Fuel. ... How Many Formula Units Of Calcium Oxide  
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Potassium Chlorate ( $\text{KClO}_3$ ) Is Heated, It Decomposes  
To Form Gases KCl And Oxygen ( $\text{O}_2$ .)  
 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$   
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