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Show Your Work, Including Proper Units, To Earn Full Credit. 1. Silver And Nitric Acid React According To The Following Balanced Equation: $3 \text{Ag(s)} + 4 \text{HNO}_3 \text{(aq)} \rightarrow 3 \text{AgNO}_3 \text{(aq)} + 2 \text{H}_2\text{O(l)} + \text{NO} \dots$ Mar 1th, 2024.

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 Chemistry 11 Stoichiometry Review Worksheet Name: ___KEY___
 33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$
 $\text{Actual yield} = 1.45\text{g}$ $\text{Theor yield} = 5.95\text{g}$
 $\text{KClO}_4 \times 1\text{mol} = 122.55\text{g}$
 $2\text{O}_2 \times 1\text{mol} = 64.00\text{g}$
 $\text{KCl} \times 1\text{mol} = 74.55\text{g}$
 $\text{Total} = 142.55\text{g}$
 $\% \text{Yield} = \frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$
 34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$
 $\text{Percent purity} = 62.8\%$
 pure FeCO_3 Apr 12th,

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