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## **Chemistry 11 Mole Fraction/Molality Worksheet Date**

1. A Solution Is Prepared By Mixing 100.0 G Of Water, H<sub>2</sub>O, And 100.0 G Of Ethanol, C<sub>2</sub>H<sub>5</sub>OH. Determine The Mole Fractions Of Each Substance. 2. The Molality Of An Aqueous Solution Of Sugar (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) Is 1.62m. Calculate The Mole Fractions Of Sugar And Water. 3. Calculate The Molality Of 25.0 Grams Of Feb 21th, 2024

## **Mole Fraction Molality Molarity - University Of Texas At ...**

Mole Fraction Sugar Of Solution  $\chi_{\text{Sugar}} = (0.292 \text{ Mol}) / (0.292 + 55.6) = .00522$  Mole Fraction Water Of Solution  $\chi_{\text{Water}} = (55.6 \text{ Mol}) / (0.292 + 55.6) = 0.995$  (or  $1 - \chi_{\text{Sugar}}$ ) Molality  $M = (.292 \text{ Mol}) / (1 \text{ Kg}) = 0.292 \text{ Mol Kg}^{-1}$  Molarity  $M = (.292 \text{ Mol}) / (1 \text{ L}) = 0.292 \text{ Mol L}^{-1}$  Jan 7th, 2024

## Calculating Molality From Molarity

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## Molarity Molality And Mole Fraction

3.mole Fraction = 0.950 4.mole Fraction = 0.0360 5.mole Fraction = 0.925  
Nsci140802exam 018 10.0points 3.3 G Of NaCl (58.4 G/mol) And 9 G Of KBr (119.8 G/mol) Were Dissolved In 40 G Of Water (18.0 G/mol). What Is The Mole Fraction Of KBr In The Solution? Nsci140805 019 10.0points How Many G Apr 5th, 2024

## K Is The Molal Freezing Point Depression And M Molality Is ...

As  $1 \text{ Atm} = 1.013 \times 10^5 \text{ Pa}$ , The Osmotic Pressure Is  $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$  Hence  $C = \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ M}^3 \text{ Pa K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$  As  $1 \text{ M}^3 = 1000 \text{ L}$ , This Corresponds To  $C = 1.5 \times 10^{-4} \text{ Mol L}^{-1}$  Feb 21th, 2024

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