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Mole Fractions Of Sugar And Water. 3. Calculate The Molality Of 25.0 Grams Of May 1th, 2024
K Is The Molal Freezing Point Depression And M Molality Is ...As 1 Atm = 1.013 × 10⁵ Pa, The Osmotic Pressure Is $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$ Hence $C = \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ M}^3 \text{ Pa K}^{-1} \text{ Mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ Mol M}^{-3}$ As 1 M³ = 1000 L, This Corresponds To $C = 1.5 \times 10$ May 3th, 2024.

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