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3 Another Example A Solution With Two Volatile Components •Modified Raoult's Law •VP T = X AP A + X BP B Example •At 20.0 OC The Vapor Pressures Of Methanol (CH 3OH) and Ethanol (C 2H 5OH) Are 95.0 And 45.0 MmHg Respectively. An Ideal Solution Contains 16.1 G Of Methanol Jan 1th, 2024

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B Raises Boiling Point Of Water From 100 O C To 100.640 C • Δ T F Lowers Freezing Point Of Water From 0o C To -2.320 C • π Raises Osmotic Pressure From 1 Atm To 28.6 Atm While The Temperature Was Changing By A Degre Feb 1th, 2024

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13 Ions In Aqueous Solutions And Colligative Properties

Ions In Aqueous Solutions And Colligative Properties MIXED REVIEW SHORT ANSWER Answer The Following Questions In The Space Provided. 1. Match The Four Compounds On The Right To Their Descriptions On The Left. B An Ionic Compound That Is Quite Soluble In Water (a) HCI C An Ionic Compound That Is Not Very Soluble In Water (b) NaNO 3 Mar 2th, 2024

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12.3 Colligative Properties - REMONDINI

Colligative Properties For Molecular Solute Vs. Ionic Solute. Example: 1.0 M Of Sugar & 1.0 M Of Salt Show Different ΔT B And ΔT F Even Though The Molality Concentrations Of The Solute Are Identical. Note That; $\Delta T = M \cdot K$ F,b Where K Is Only A Function Of The Solvent !!! Reason For The Difference In Colligative Properties Must Be Based On ... Feb 1th, 2024

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