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= $k_1[D]_1 X[B]_1 Y$ $k_2[D]_2 X[B]_2 Y$ Equation 5 In This Equation k Cancels Out. $[D]_1 = [D]_2$ Because We Are Starting Off With The Same Initial Concentration Of A In

Each Trial. The Order Of Species D Which Is X Is Also Constant. Canceling Terms We Have Left: $\text{Rate}_1 = k_1[B]_1$ $\text{Rate}_2 = k_2[B]_2$ Equation 6 Consider The Following Chemical Reaction: Jan 6th, 2024

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2nd Cl 2to FormHCl. 7 Chemistry-I. Figure 1.1: Plot Of Concentration Against T For

A Reaction $H_2 + Cl_2 \rightarrow 2HCl$ The Rate Of This Reaction Does Not Depend On Concentration. 1.6 First Order Kinetics A Reaction Of The first Order Is Represented As $X \rightarrow Y$ Where X Is The Reactant And Y The Product. The Rate Of The Reaction Will Be Directly Feb 3th, 2024

Chemical Kinetics - Duke University

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B. Transfer Translational KE To Vibrational Energy To Weaken/break Reactant Bonds
Molecules At A Given Temperature Possess A KE Distribution: Number Of Molecules
KE $E \geq E_A$, Activation Energy - Energy Barrier That Molecules Have To Surmount In
Order To React. Energy Is Needed To Break Reactant Bonds (endothermic Process).
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Chemical Identity - What Exactly Is Reacting? Usually Ions Of Opposite Charge React
Very Rapidly. Also, The More Bonds Between Reacting Atoms In A Molecule, The
Slower The Reaction Rate. Why? More Energy Is Required To Separate The Molecule
Into Its "bits". Substances With Strong Bonds (larger Bond Energies) Will React
Much More Slowly. May 6th, 2024

Chapter 14 Chemical Kinetics

3 Chemical Kinetics Factors That Affect Reaction Rates • Physical State Of The
Reactants In Order To React, Molecules Must Come In Contact With Each Other. If
The Reaction Is Happening Between A Solid And A Liquid It Will React Only On The

Surface. The More Homogeneous The Mixture Of Reactants, The Faster The Molecules Can React. May 17th, 2024

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